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~~class-12, unit-7 Chemistry~~

~~Review Sheet Unit 7~~

are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes.

1. Calculate the molar mass for ethane, C_2H_6 .
 C_2H_6 $2(12.0 \text{ g}) + 6(1.0 \text{ g}) = 30.0 \text{ g}$

2. Convert 45 g Se to moles of Se. ? mol Se
 $= 45 \text{ g Se} \cdot \frac{1 \text{ mol Se}}{79.0 \text{ g Se}} = 0.57 \text{ mol Se}$

3.

~~Review Sheet: Unit 7 Name~~

~~Vocabulary: Fill in the blanks ...~~

are representative of each type of problem worked in this unit. For more practice

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on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane, C_2H_6 . 2. Convert 45 g Se to moles of Se. 3. Convert 4.77×10^{24} molecules of SO_2 to grams.

~~Review Sheet: Unit 7 Name~~
~~Vocabulary: Fill in the blanks ...~~

UNIT 7 REVIEW SHEET -
EQUILIBRIUM • EQUILIBRIUM is the point at which the rate of the forward reaction is equal to the rate of the reverse reaction in a reversible process. o
Reaction is dynamic - still occurring - even though to

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the eye, it looks like nothing is changing at equilibrium. o

Observable/measurable quantities (concentration, pressure, color) are constant at equilibrium. o The amounts of products and reactants are typically not equal at equilibrium!

~~UNIT 7 REVIEW SHEET.pdf~~
~~UNIT 7 REVIEW SHEET \u2013~~

...

CHEMISTRY CP Name: _____

Per: _____ TEST: _____ Unit
7 Review Sheet:

Thermochemistry KEY

Potential Energy Diagrams:

Choose the appropriate letter. 1. What is the ΔH of the forward reaction? f 2.

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What is the activation energy of the forward reaction? b 3. What is the H of the reverse reaction? f 4.

~~Chemistry Unit 7 Review Sheet KEY - OAK PARK USD~~

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Review Sheet: Unit 5

Name_____ CHEMISTRY: A Study of Matter © 2004, GPB 5.24a
KEY I. For each of the following statements, write "I" for ionic, "C" for covalent

~~CHEMISTRY: A Study of Matter~~
Review Sheet for Chemistry
Unit 5 December 11, 2011 31.
Balance the following equations by putting the proper coefficients wherever they are needed. a) $\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$ b) $\text{Al} + \text{H}_2\text{O} \rightarrow \text{Al}(\text{OH})_3 + \text{H}_2$ c) $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ d) C_5H_{12}

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+ O₂ CO₂ + H₂O e) Na + H₂O
H₂ + NaOH f) NaClO₃ NaCl
+ O₂ g ...

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